



Universidade Estadual do Norte Fluminense Darcy Ribeiro



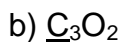
CURSO: LICENCIATURA EM QUÍMICA – EXERCÍCIO PROGRAMÁTICO 7

POLOS: Nova Friburgo, Paracambi, São Fidelis e São Francisco

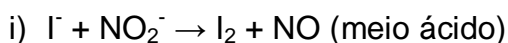
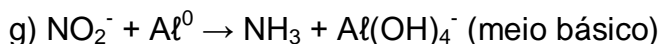
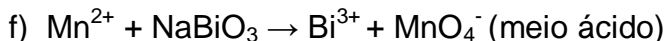
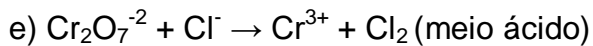
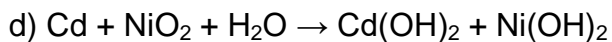
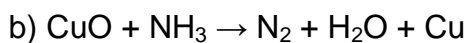
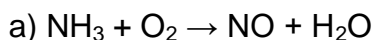
DISCIPLINA: QUÍMICA GERAL III

PERÍODO: 2017-1

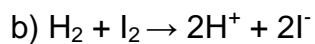
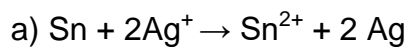
Exercício 1: Dê o número de oxidação dos elementos sublinhados em cada sentença:



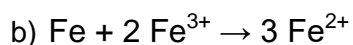
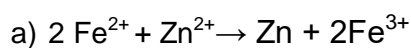
Exercício 2: Faça o balanceamento das seguintes questões:



Exercício 3: Faça o diagrama esquemático de uma célula galvânica, indicando o anodo, o catodo, o fluxo de elétrons, cátions e ânions, e calcule o potencial padrão das seguintes células:



Exercício 4: Calcule o potencial padrão das seguintes células galvânicas e diga se a reação ocorrerá de maneira espontânea.



Exercício 5: Calcule o potencial gerado pela pilha descrita abaixo quando $[\text{Al}^{3+}] = 4,0 \times 10^{-3} \text{ mol.L}^{-1}$ e $[\text{I}^-] = 0,010 \text{ mol.L}^{-1}$

